

Fill in the table below

concept	Describe how the electron is involved	Give <u>1</u> example how the concept is applied:
Cations	Less electrons than protons. Since there are more protons, the atom has an overall positive charge. Atom is trying to obtain a noble gas configuration. Metals lose electrons and form cations.	Positively charged atom, bonds ionically or covalently in-order to become stable. Positively charged atom is attracted to a negatively charged atom. Covalent - Shares electrons with valence electrons of another atom to obtain the stable octet. Electron clouds overlap during sharing.
Anions		
Stability		
Atomic orbitals, (spdf)		
Families/groups		
VSEPR		
Covalent bonding		
Ionic bonding		
Octet		
London dispersion force		

concept	Describe how the electron is involved	Give <u>1</u> example how the concept is applied:
Electronegativity		
Electron affinity		
Oxidation number		
Metallic bonding		
Polarity		
Hydrogen bonding		
Ionization energy		
Electron configuration		
Conductivity		
Dipole dipole force		
oxidation		
reduction		